

a1 $V_{10}O_{28}^{6-}$ and $Mo_7O_{24}^{6-}$, monocarboxylates such as acetate, dicarboxylates such as oxalate, and alkyl sulphonates such as laurylsulphonate.

Page 1, lines 22-26, delete current paragraph and insert therefor:

It should be noted that a variety of terms are used to describe the material which is referred to herein as an anionic clay. Hydrotalcite-like and layered double hydroxide are interchangeably used by those skilled in the art. Herein, the materials are referred to as *a2* anionic clays, comprising within that term hydrotalcite-like and layered double hydroxide materials.

Page 2, lines 3-10, delete current paragraph and insert therefor:

Recently, two major reviews of anionic clay chemistry were published in which the synthesis methods available for anionic clay synthesis have been summarized, F. Cavani et al., "Hydrotalcite-type anionic clays: Preparation, Properties and Applications," Catalysis Today, 11 (1991) Elsevier Science Publishers B. V. Amsterdam.

a3 J P Bessee and others, "Anionic clays: trends in pillary chemistry, its synthesis and microporous solids" (1992), 2, 108, editors: M.I. Occelli, H.E. Robson, Van Nostrand Reinhold, N.Y.

Page 2, lines 12-19, delete current paragraph and insert therefor:

In these reviews the authors state that a characteristic of anionic clays is that mild calcinations at 500°C results in the formation of a disordered MgO-like product. The disordered MgO-like product is distinguishable from spinel (which results upon severe *a4* calcination) and from anionic clays. Herein, the disordered MgO-like materials are referred to as Mg-Al solid solutions. Furthermore, these Mg-Al solid solutions contain a well-known memory effect whereby the exposure to water of such calcined materials results in the reformation of the anionic clay structure.

Page 3, lines 10-17, delete current paragraph and insert therefor:

In US Patent No. 3,796,792 by Miyata entitled "Composite Metal Hydroxides" a range of materials is prepared into which an extensive range of cations is incorporated, including Sc, La, Th, In, etc. In the examples given solutions of the divalent and trivalent cations are prepared and mixed with base to cause co-precipitation. The resulting products are filtered, washed with water, and dried at 80°C. Example 1 refers to Mg and Al and Example 2 to Mg and Bi. Other examples are given, and in each case soluble salts are used to make solutions prior to precipitation of the anionic clay at high pH.

②5 Page 3, lines 19-27, delete current paragraph and insert therefor:

In US Patent No. 3,879,523 by Miyata entitled "Composite Metal Hydroxides" also a large number of preparation examples is outlined. The underlying chemistry, however, is again based on the co-precipitation of soluble salts followed by washing and drying. It is important to emphasize that washing is a necessary part of such preparations, because to create a basic environment for co-precipitation of the metal ions a basic solution is needed and this is provided by NaOH/Na₂CO₃ solutions. Residual sodium, for example, can have a significant deleterious effect on the subsequent performance of the product as a catalyst or oxide support.

②6 Page 4, line 1, delete current paragraph and insert therefor:

②7 In US Patent No. 3879525 (Miyata) very similar procedures are again described.

②8 Page 4, lines 3-10, delete current paragraph and insert therefor:

In US Patent No. 4,351,814 to Miyata et al. a method for making fibrous hydrotalcite is described. Such materials differ in structure from the normal plate-like morphology. The synthesis again involves soluble salts. For example, an aqueous solution of a mixture of MgCl₂ and CaCl₂ is prepared and suitably aged. From this a needle-like product Mg₂(OH)₃Cl·4H₂O precipitates. A separate solution of sodium aluminate is then reacted in

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an autoclave with the solid $Mg_2(OH)_3Cl \cdot 4H_2O$ and the product is again filtered, washed with water, and dried.

Page 4, lines 12-16, delete current paragraph and insert therefor:

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In US Patent No. 4,458,026 to Reichle, in which heat-treated anionic clays are described as catalysts for aldol condensation reactions, again use is made of magnesium and aluminum nitrate salt solutions. Such solutions being added to a second solution of NaOH and Na_2CO_3 . After precipitation the slurry is filtered and washed twice with distilled water before drying at 125°C.

Page 4, lines 18-25, delete current paragraph and insert therefor:

a 10
In US Patent No. 4,656,156 to Misra the preparation of a novel absorbent based on mixing activated alumina and hydrotalcite is described. The hydrotalcite is made by reacting activated MgO (prepared by activating a magnesium compound such as magnesium carbonate or magnesium hydroxide) with aqueous solutions containing aluminate, carbonate and hydroxyl ions. As an example the solution is made from NaOH Na_2CO_3 and Al_2O_3 . In particular, the synthesis involves the use of industrial Bayer liquor as the source of Al. The resulting products are washed and filtered before drying at 105°C.

Page 4, line 27 - page 5, line 1, delete current paragraph and insert therefor:

a 11
In US Patent No. 4,904,457 to Misra a method is described for producing hydrotalcites in high yield by reacting activated magnesia with an aqueous solution containing aluminate, carbonate, and hydroxyl ions.

Page 5, line 3, delete current paragraph and insert therefor:

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The methodology is repeated in US Patent No. 4,656,156.

Page 5, lines 5-12, delete current paragraph and insert therefor:

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In US Patent No. 5,507,980 to Kelkar et al. a process is described for making novel catalysts, catalyst supports, and absorbers comprising synthetic hydrotalcite-like binders. The

synthesis of the typical sheet hydrotalcite involves reacting pseudo-boehmite to which acetic acid has been added to peptize the pseudo-boehmite. This is then mixed with magnesia.

a 10 More importantly, the patent summary states clearly that the invention uses mono carboxylic organic acids such as formic, propionic and isobutyric acid. In this patent the conventional approaches to preparing hydrotalcite are presented.

Page 5, lines 14-18, delete current paragraph and insert therefor:

In US Patent No. 5,439,861 a process is disclosed for preparing a catalyst for synthesis gas production based on hydrotalcite. The method of preparation is again based, on the co-
a 10 precipitation of soluble salts by mixing with base, for example, by the addition of a solution of RhCl_3 , $\text{Mg}(\text{NO}_3)_2$ and $\text{Al}(\text{NO}_3)_3$ to a solution of Na_2CO_3 and NaOH .

Page 5, lines 20-22, delete current paragraph and insert therefor:

Also in US Patent No. 5,399,537 to Bhattacharyya in the preparation of nickel-containing catalysts based on hydrotalcite use is made of the co-precipitation of soluble magnesium and aluminum salts.

Page 5, lines 24-28, delete current paragraph and insert therefor:

In US Patent No. 5,591,418 to Bhattacharyya a catalyst for removing sulfur oxides or
a 10 nitrogen oxides from a gaseous mixture is made by calcining an anionic clay, the anionic clay having been prepared by co-precipitation of a solution of $\text{Mg}(\text{NO}_3)_2$, $\text{Al}(\text{NO}_3)_3$ and $\text{Ce}(\text{NO}_3)_3$. The product again is filtered and repeatedly washed with de-ionized water.

Page 6, lines 1-7, delete current paragraph and insert therefor:

i In US Patent No. 5,114,898/WO 9110505 Pinnavaia et al. describe layered double hydroxide sorbents for the removal of sulfur oxide(s) from flue gases, which layered double hydroxide is prepared by reacting a solution of Al and Mg nitrates or chlorides with a solution of NaOH and Na_2CO_3 . In US Patent No. 5,079,203/WO 9118670 layered double hydroxides

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intercalated with polyoxo anions are described, with the parent clay being made by co-
precipitation techniques.

Page 6, lines 9-14, delete current paragraph and insert therefor:

a 18
In US Patent No. 5,578,286 in the name of Alcoa a process for the preparation of
meixnerite is described. The meixnerite may be contacted with a dicarboxylate or
polycarboxylate anion to form a hydrotalcite-like material. In comparative examples 1-3
hydromagnesite is contacted with aluminum trihydrate in a CO₂ atmosphere, greater than 30
atmospheres. No hydrotalcite was obtained in these examples.

Page 6, lines 16-19, delete current paragraph and insert therefor:

a 19
In US Patent No. 5,514,316 a method for the preparation of meixnerite is described
using magnesium oxide and transition alumina. For comparative purposes aluminum
trihydrate was used in combination with magnesium oxide. It was indicated that this method
did not work as well as the transition alumina.

Page 6, lines 21-22, delete current paragraph and insert therefor:

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US Patent Nos. 4,454,244 and 4,843,168 describe the use of pillaring anions in
anionic clays.

Page 6, lines 24-28, delete current paragraph and insert therefor:

a 20
In US Patent Nos. 4,946,581 and 4,952,382 to van Broekhoven co-precipitation of
soluble salts such as Mg(NO₃)₂ and Al(NO₃)₃ with, and without the incorporation of rare
earth salts was used for the preparation of anionic clays as catalyst components and additives.

A variety of anions and di- and tri-valent cations are described.

Page 7, lines 9-26, delete current paragraph and insert therefor:

a 22
Because of this wide variety of large-scale commercial applications for these
materials, new processes utilizing alternative inexpensive raw materials and which can be
carried out in continuous mode are needed to provide more cost-effective and

environmentally compatible processes for making anionic clays. In particular, from the prior art described above one can conclude that the preparation process can be improved in the following ways: the use of cheaper sources of reactants, processes for easier handling of the reactants, so that there is no need for washing or filtration, eliminating the filtration problems associated with these fine-particled materials, the avoidance of alkali metals (which can be particularly disadvantageous for certain catalytic applications): In prior art preparations *a 27* organic acids were used to peptize alumina. The use of organic acids is expensive and introduces an additional step in the synthesis process and is therefore not cost-effective. Further, in drying or calcining the anionic clay prepared by prior art processes gaseous emissions of nitrogen oxides, halogens, sulfur oxides, etc. are encountered which cause environmental pollution problems. Moreover, none of the preparation methods described in the prior art provide continuous processes for the preparation of anionic clays.

Page 8, lines 3-10, delete current paragraph and insert therefor:

Our invention includes a process for producing anionic clays using raw materials which are inexpensive and utilizing such raw materials in a simple process which is extremely suitable to be carried out in continuous mode. The process involves reacting *a 28* mixtures with or without stirring in water at ambient or elevated temperature at atmospheric or elevated pressure. Such continuous processes can be operated in standard industrial equipment. More specifically, there is no need for washing or filtering, and a wide range of ratios of Mg/Al in the reaction product is possible.

Page 8, lines 11-26, delete current paragraph and insert therefor:

In the process according to the invention an aluminum source and a magnesium *a 29* source, for instance magnesium oxide or brucite, are reacted in aqueous suspension to obtain an anionic clay. The aluminum source comprises two types of aluminum-containing compounds, for instance alumina trihydrate (such as gibbsite, bayerite or nordstrandite) and

thermally treated forms thereof. The reaction is carried out at ambient or elevated temperature and ambient or elevated pressure and the reaction mixture results in the direct formation of an anionic clay which can be obtained by simply drying the slurry continuously retrieved from the reactor. The powder X-ray diffraction pattern (PXRD) suggests that the product is comparable to anionic clays made by other standard (batch) methods. The physical and chemical properties of the product are also comparable to those anionic clays made by the other conventional methods. The overall process of this invention is very flexible, enabling a wide variety of anionic clay compositions and anionic clay-like materials involving for example carbonate, hydroxide and other anions to be prepared in an economically and environmental-friendly manner.

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Page 9, line 18 - page 10, line 5, delete current paragraph and insert therefor:

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This invention involves the preparation of anionic clays. In particular it describes a process for the preparation of an anionic clay wherein an aluminum source and a magnesium source are reacted in aqueous suspension to obtain an anionic clay, the aluminum source comprising two types of aluminum-containing compounds wherein one type of aluminum-containing compound is aluminum trihydrate or its thermally treated form. The magnesium source may be composed of a solution of a magnesium salt, a solid magnesium-bearing compound or a mixture of the two. Reaction between the Mg source and aluminum source results in the direct formation of an anionic clay. The reaction takes place at room temperature or higher. At temperatures higher than 100°C, the reaction is preferably carried out under autogeneous conditions. In the method according to the invention carbonate, hydroxyl, or other anions or mixtures thereof, either provided within the reaction medium for example by feeding a soluble salt to the reactor or absorbed during the synthesis from the atmosphere, are incorporated into the interlayer region as the necessary charge-balancing anion.

Page 11, lines 3-7, delete current paragraph and insert therefor:

Wherein m and n have a value such that m/n = 1 to 10, preferably 1 to 6, and b has a value in the range of from 0 to 10, generally a value of 2 to 6 and often a value of about 4.

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~~any~~ $X_{n/z}^{z-}$ may be CO_3^{2-} , OH^- and any other anion normally present in the interlayers of anionic clays. It is more preferred that m/n should have a value of 2 to 4; more particularly a value close to 3.

Page 11, lines 9-14, delete current paragraph and insert therefor:

Since the process disclosed herein does not require washing of the product or filtering, there is no filtrate waste or gaseous emissions (e.g. from acid decomposition), making the process particularly environmental-friendly and more suited to the environmental constraints which are increasingly imposed on commercial operations. The product can be spray dried directly to form microspheres or can be extruded to form shaped bodies.

Page 11, line 16 - page 12, line 27, delete current paragraph and insert therefor:

Aluminum source

The present invention includes the use of two types of aluminum-containing compounds wherein one type of aluminum-containing compound is crystalline aluminum trihydrate (ATH) or its thermally treated form. An example of aluminum trihydrate is gibbsite (for instance provided by Reynolds Aluminum Company RH-20® or JM Huber *a 27* Micral ® grades). Also BOC (Bauxite Ore Concentrate), bayerite and nordstrandite are suitable aluminum trihydrates. BOC is the cheapest alumina source. The alumina trihydrate is preferred to have a small particle size. Thermally treated forms of gibbsite can also be used. The thermally treated (calcined) aluminum trihydrate is readily obtained by thermally treating aluminum trihydrate (gibbsite) at a temperature ranging from 100 to 800°C for 15 minutes to 24 hours. In any event, the calcining temperature and time for obtaining calcined aluminum trihydrate should be sufficient to cause a measurable increase of the surface area in

view of the surface area of the gibbsite as produced by the Bayer process which is generally between 30 and 50 m²/g. It should be noted that within the concept of this invention flash calcined alumina is also considered to be a thermally treated form of aluminum trihydrate, although generally it is considered a very specific alumina. Flash calcined alumina is obtained by treating aluminum trihydrate at temperatures between 800-1000°C for very short periods of time in special industrial equipment, as is described in US Patent Nos. 4,051,072 and 3,222,129. When using aluminum trihydrate other aluminum-containing compounds such as oxides and hydroxides of aluminum, (e.g. sols, thermally treated aluminum trihydrate including flash calcined alumina, gels, pseudo-boehmite, boehmite) aluminum salts such as aluminum nitrate, aluminum chloride, aluminum chlorohydrate and sodium aluminate are added as the second type of aluminum-containing compound. The other aluminum-containing compounds may be soluble or insoluble in water and may be added to the aluminum trihydrate or it may be added to the reactor separately as a solid, a solution or as a suspension. When using a thermally treated aluminum trihydrate also other aluminum-containing compounds are added such as the ones described above and of course aluminum trihydrate and other thermally treated forms thereof. The other aluminum sources may be added to the thermally treated aluminum trihydrate or it may be added to the reactor separately as a solid, a solution or as a suspension. Preferably, the aluminum source is added to the reactor in the form of a slurry. In particular we emphasize that there is no need to use a peptizable alumina source (gibbsite is not peptizable) and as a result no need to add either mineral or organic acid to vary the pH of the mixture.

Page 13, lines 14-19, delete current paragraph and insert therefor:

Because of its simplicity, this process is particularly suitable to be carried out in a

a29 continuous mode. Thereto an aluminum source and a magnesium source are fed to a reactor and reacted in aqueous suspension to obtain an anionic clay. In the case of a batch process

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the aluminum source and magnesium source are added to the reactor and reacted in aqueous suspension to obtain an anionic clay.

Page 13, line 20 - page 14, line 3, delete current paragraph and insert therefor:

Within the context of this invention a reactor is considered to be any confined zone in which the reaction between the aluminum source and magnesium source takes place. The reactor may be equipped with stirrers, baffles, etc. to ensure homogeneous mixing of the reactants. The reaction can take place with or without stirring, at ambient or at elevated temperature and at atmospheric or elevated pressure. Usually, a temperature between 0 and 100°C is used at or above atmospheric pressure. It is preferred to carry out the process at temperatures above 50°C rather than at room temperature, because this results in anionic clays with sharper peaks in the x-ray diffraction pattern than anionic clays obtained at room temperature. The reactor may be heated by any heating source such as a furnace, microwave, infrared sources, heating jackets (either electrical or with a heating fluid), lamps, etc.

a 30
Page 14, lines 5-16, delete current paragraph and insert therefor:

The aqueous suspension in the reactor may be obtained by either feeding slurries of the starting materials, either combined or separate, to the reactor or adding magnesium source to a slurry of alumina source or vice versa and feeding the resulting slurry to the reactor. It is possible to treat, for instance the aluminum source slurry at elevated temperature and then add either the Mg source per se, or add the Mg source in a slurry or solution either to the reactor or the aluminum source slurry. Given particular facilities which might be available, the process can be conducted hydrothermally. This is particularly advantageous, because it is faster and a higher conversion is obtained. There is no need to wash or filter the product, as unwanted ions (e.g. sodium, ammonium, chloride, sulfate) which are frequently encountered when using other preparation methods, are absent in the product.

Page 14, lines 17-23, delete current paragraph and insert therefor:

In a further embodiment of the invention, the process is conducted in a multi-step process, e.g. a slurry of aluminum source and magnesium source is treated thermally in a first reactor at a mild temperature, followed by a hydrothermal treatment in a second reactor. If *a 31* desired a preformed anionic clay may be added to the reactor. The preformed clay may be recycled anionic clay from the reaction mixture or anionic clay made separately by the process according to the invention or any other process.

Page 14, line 25 - page 15, line 1, delete current paragraph and insert therefor:

If desired, organic or inorganic acids and bases, for example for control of the pH, may be added to the reactor or added to either the magnesium source or the aluminum source *a 33* before they are added to the reactor. An example of a preferred pH modifier is an ammonium base, because upon drying no deleterious cations remain in the anionic clay.

Page 15, lines 3-10, delete current paragraph and insert therefor:

If desired, the anionic clay prepared by the process according to the invention may be subjected to ion exchange. Upon ion exchange the interlayer charge-balancing anions are replaced with other anions. The other anions are the ones commonly present in anionic clays *a 34* and include pillaring anions such as $V_{10}Q_{28}^{-6}$, $Mo_7O_{24}^{6-}$, $PW_{12}O_{40}^{3-}$, $B(OH)_4^-$, $B_4O_5(OH)_4^{2-}$, HBO_4^{2-} , $HGaO_3^{2-}$, CrO_4^{2-} . Examples of suitable pillaring anions are given in US Patent No. 4774212 which is included by reference for this purpose. The ion exchange can be conducted before or after drying the anionic clay formed in the slurry.

Page 15, line 17 - page 16, line 1, delete current paragraph and insert therefor:

For some applications it is desirable to have additives, both metals and non-metals, *a 35* such as rare earth metals, Si, P, B, group VI, group VII, alkaline earth (for instance Ca and Ba) and/or transition metals (for example Mn, Fe, Ti, Zr, Cu, Ni, Zn, Mo, Sn), present. The

metals and non-metals can easily be deposited on the anionic clay or the solid solution according to the invention or they can be added either to the alumina source or magnesium source which are added combined to the reactor or separately. The metals and non-metals can also be added to the aqueous suspension in which the reaction takes place. Suitable sources of metals or non-metals are oxides, halides or any other salt such as chlorides, nitrates etc. In the case of a multi-step process the metals and non-metals may be added in any of the steps. It can be especially advantageous for controlling the distribution of the metals and non-metals in the anionic clay.

Page 16, lines 19-21, delete current paragraph and insert therefor:

This comparative example illustrates the co-precipitation method where Mg and Al salt solutions are added to a solution of base. (US Patent No. 3 979 523, Assignee Kyowa Chemical Industry, Japan)

Page 19, lines 4-11, delete current paragraph and insert therefor:

16.27 g gibbsite and a sample of 5.13 g gibbsite previously calcined at 400°C were slurried in 400 g deionized water. To this a slurry of 27.2 g MgO powder in 170 g deionized water was added and mixed in blender for ten minutes. The weight ratio of gibbsite to calcined gibbsite was 70:30 and the overall Mg:Al ratio in the slurry was 2.3. The pH was adjusted to 9.94 by the addition of ammonium hydroxide solution. Final slurry solids = 7.0 wt%. The slurry was aged at 120°C for 18 hours and the product dried at 110°C. See attached Figure 4.

Page 19, lines 15-22, delete current paragraph and insert therefor:

16.27 g gibbsite and a sample of 5.13 g gibbsite previously calcined at 400°C were slurried in 400 g deionized water. To this a slurry of 27.2 g MgO powder in 170 g deionized water was added and mixed in blender for ten minutes. The weight ratio of gibbsite to calcined gibbsite was 70:30 and the overall Mg:Al ratio in the slurry was 2.3. The pH was

adjusted to 9.94 by the addition of ammonium hydroxide solution. Final slurry solids = 7.0 wt%. The slurry was aged at 200 psi (ca. 198°C) in a microwave over for 60 minutes. The product was dried at 110°C.

Page 19, line 26 - page 20, 4, delete current paragraph and insert therefor:

a 38 16.27 g gibbsite and a sample of 5.75 g flash calcined gibbsite CP-1.5 ® were slurried in 400 g deionized water. To this a slurry of 27.2 g MgO powder in 170 g deionized water was added and mixed in blender for ten minutes. The weight ratio of gibbsite to CP-1.5 was 70:30 and the overall Mg:Al ratio in the slurry was 2.3. The pH was adjusted to 9.87 by the addition of ammonium hydroxide solution. Final slurry solids = 7.0 wt%. The slurry was aged at 120°C for 18 h and the product dried at 110°C. See attached Figure 5.

Page 20, lines 8-15, delete current paragraph and insert therefor:

a 39 16.27 g gibbsite and a sample of 5.75 g flash calcined gibbsite CP-1.5 ® were slurried in 400 g deionized water. To this a slurry of 27.2 g MgO powder in 170 g deionized water was added and mixed in blender for ten minutes. The weight ratio of gibbsite to CP-1.5 was 70:30 and the overall Mg:Al ratio in the slurry was 2.3. The pH was adjusted to 9.87 by the addition of ammonium hydroxide solution. Final slurry solids = 7.0 wt%. The slurry was treated at 200 psi (ca. 198°C) in a microwave oven for 60 minutes. The product was dried at 110°C.

Page 20, lines 19-25, delete current paragraph and insert therefor:

a 40 16.27 g gibbsite and a sample of 6.09 g Catapal A ® were slurried in 400 g deionized water. To this a slurry of 27.2 g MgO powder in 170 g deionized water was added and mixed in blender for ten minutes. The weight ratio of gibbsite to Catapal was 70:30 and the overall Mg:Al ratio in the slurry was 2.3. The pH was adjusted to 9.96 by the addition of ammonium

a 40 hydroxide solution. Final slurry solids = 7.0 wt%. The slurry was aged at 120°C for 18 hours and the product dried at 110°C. See attached figure 6.

a 41 Page 21, lines 4-10, delete current paragraph and insert therefor:

16.27 g gibbsite and a sample of 6.09 g Catapal A ® were slurried in 400 g deionized water. To this a slurry of 27.2 g MgO powder in 170 g deionized water was added and mixed in blender for ten minutes. The weight ratio of gibbsite to catapal was 70:30 and the overall Mg:Al ratio in the slurry was 2.3. The pH was adjusted to 9.96 by the addition of ammonium hydroxide solution. Final slurry solids = 7.0 wt%. The slurry was treated at 200 psi (ca. 198°C) in a microwave oven for 60 minutes. The product was dried at 110°C.

IN THE CLAIMS:

Please cancel claims 1-3 and 6-8 without prejudice to or disclaimer of the subject matter contained therein.

Please replace claims 4-5 and 9-15 as follows:

a 42 4. (Amended) The process according to claim 23, wherein the aluminum source comprises aluminum trihydrate and thermally treated aluminum trihydrate.

a 43 5. (Amended) The process according to claim 23, wherein the magnesium source is at least one selected from the group consisting of MgO, Mg(OH)₂ and MgCO₃.

a 44 9. (Amended) The process according to claim 23, wherein metals or non-metals are fed to the reactor.

a 45 10. (Amended) The process according to claim 9, wherein the metals or non-metals are added to an aluminum source slurry.

a 46 11. (Amended) The process according to claim 9, wherein the metals or non-metals are added to a magnesium source slurry.

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12. (Amended) The process according to claim 23, wherein the anionic clay is subjected to an ion-exchange treatment.

13. (Amended) The process according to claim 23, wherein the anionic clay is ion exchanged with pillaring anions such as $V_{10}O_{28}^{6-}$ and $Mo_7O_{24}^{6-}$.

14. (Amended) The process according to claim 23, wherein metals or non-metals are deposited on the anionic clay.

15. (Amended) The process for the preparation of a Al-Mg-containing solid solution and/or spinel, comprising subjecting an anionic clay obtained by the process of claim 23 is subjected to a heat-treatment at a temperature between 300 and 1200°C.

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Please add new claims 16-25 as follows:

--16. A process for the preparation of anionic clays, comprising reacting an aluminum source and a magnesium source in aqueous suspension to obtain an anionic clay, the aluminum source comprising two types of aluminum-containing compounds, wherein one type of aluminum-containing compound is thermally treated aluminum trihydrate.--

--17. The process according to claim 16, wherein the thermally treated aluminum trihydrate is thermally treated gibbsite.--

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--18. The process according to claim 16, wherein one of the two types of aluminum-containing compounds is thermally treated gibbsite, and the other one of the two types of aluminum-containing compounds is gibbsite.--

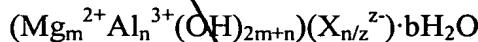
--19. The process according to claim 16, wherein the reaction takes place at a temperature between 0 and 100°C and at or above atmospheric pressure.--

--20. The process according to claims 19, wherein the reaction takes place at a temperature above 50°C and at or above atmospheric pressure.--

--21. The process according to claim 16, wherein the reaction is conducted at a temperature above 100°C and at a pressure above atmospheric pressure.--

--22. The process according to claim 16, wherein the magnesium source is at least one selected from the group consisting of MgO, Mg(OH)₂ and MgCO₃--

--23. A process for the preparation of anionic clays corresponding to the general formula



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wherein m and n have a value such that m/n = 1 to 10, b has a value in the range of from 0 to 10, and X_{n/z}^{z-} may be CO₃²⁻, OH⁻, or any other anion present in the interlayers of the anionic clays, which process comprises reacting an aluminum source and a magnesium source in aqueous suspension at a temperature above 100°C and at a pressure above atmospheric pressure to obtain an anionic clay, the aluminum source comprising two types of aluminum-containing compounds, wherein the first type of aluminum-containing compound is either aluminum trihydrate or its thermally treated form and wherein

(a) when the first type of aluminum-containing compound is aluminum trihydrate, the second type of aluminum-containing compound is selected from the group consisting of aluminum sols, thermally treated aluminum trihydrate, aluminum gels, pseudoboehmite, boehmite, aluminum nitrate, aluminum chloride and aluminum chlorohydrate, and

(b) when the first type of aluminum-containing compound is thermally treated aluminum trihydrate, the second type of aluminum-containing compound is selected from the group consisting of other forms of thermally treated aluminum trihydrate, aluminum trihydrate, aluminum sols, aluminum gels, pseudoboehmite, boehmite, aluminum nitrate, aluminum chloride and aluminum chlorohydrate.--

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